

CHART - 5 > Name or state the following

1. A liquid oxidizing agent and a gaseous reducing agent.
2. A salt or a gas which undergoes a photochemical reaction.
3. The salt formed when zinc reacts with caustic potash.
4. The group number of the element having atomic number 12.
5. The subatomic particle which carries a unit negative charge and has negligible mass.
6. A metal other than iron and copper which shows variable valency.
7. Two neutral gases one of which is oxygen which combines to give a coloured acidic gas.
8. The acid responsible for slight acidity in natural rain water.
9. A gas which is combustible but a non-supporter of combustion.
10. The method used to separate colouring matter in ink.
11. An acidic gas released both during respiration and burning.
12. The gas liberated when ammonium dichromate undergoes thermal decomposition.
13. A gas responsible for melting of ice-caps.
14. A salt whose solubility decreases with increase in temperature of water.
16. A metal which burns with a brilliant yellow flame in oxygen.
17. A hygroscopic liquid which acts as a drying and a dehydrating agent.
18. A liquid which in the presence of a catalyst, evolves oxygen.
19. The chemical other than chlorofluorocarbon, responsible for ozone depletion & global warming.
20. An exothermic reaction involving two neutral gaseous reactants.
21. An amphoteric oxide other than zinc oxide and lead monoxide.
22. A non-metal in group 17 [VIIA] of the periodic table, which is a solid at ordinary temperatures.
23. The salt formed when aluminium reacts with conc. sodium hydroxide solution.
24. An example of a 'mixed acid anhydride'.
25. The process involving hydrogen used in the manufacture of vegetable fat.
26. The type of oxide formed by the element in period 3 and group 13 [IIIA].
27. The most abundant element in the earth's crust.
28. An element other than hydrogen and chlorine which exists in the isotopic form.
29. A lead salt which evolves reddish brown fumes on thermal decomposition.
30. A deliquescent substance also used in the softening of hard water.
31. The more electronegative element from the elements oxygen and sulphur present in group 16 [VIA].
32. The law which states that the product of the volume and pressure of a given mass of a dry gas is constant provided temperature remains constant.
33. The most reactive and the least reactive metal from the metals Na, Al, Cu, Ag.
34. The element in group 17 [VIIA] of the periodic table which is a liquid at ordinary temperatures.
35. The period which contains 8 elements including the non-metal sulphur.
36. The element in group 17 [VIIA] of the periodic table which bleaches vegetables dyes by oxidation.

37. A metal other than platinum which does not form an oxide.
38. The most reactive element in group 16 [VIA] of the periodic table.
39. The product obtained when nitrogen reacts with a neutral gas and the reaction is endothermic.
40. A nitride of a divalent metal.
41. The displaced product when zinc reacts with copper sulphate solution.
42. The colourless liquid obtained when the first element of group 14 [IVA], reacts with the second element of group 16 [VIA] of the periodic table.
43. The least reactive element in group 15 [VA] of the periodic table.
44. The state of matter from solids, liquids & gases, whose inter-particle attraction is maximum & energy possessed by particles are least.
45. The term which represents the change from gaseous state to liquid state, without any fall in temperature.
46. The gaseous state or vapour formed, when a solid directly changes to gaseous state, without changing into liquid state.
47. The Law which states that -
 - a] The product of the volume and pressure of a given mass of dry gas is constant -temperature remaining constant.
 - b] The volume of a given mass of any gas is directly proportional to its absolute temperature-pressure remaining constant.
 - c] In a chemical reaction, the total mass of the reacting substances is equal to the total mass of the products - masses measured under similar conditions.
48. The temperature scale with absolute zero as its starting point.
49. An element which is inactive and present in traces in the atmosphere.
50. A secondary pollutant from the primary pollutant nitric oxide.
51. A homogeneous mixture of -
 - a] two liquids b] liquid & gas- in which the liquid component is water in each case.
52. The method used for separation of a] two miscible liquids b] two immiscible liquids.
53. A a] positive radical b] negative radical - both containing the element 'H'.
54. A metal which shows variable valency of a] +1 & +2 b] +2 & +3 c] +2 & +4.
55. The source of inorganic material in sewage waste water.
56. A combustible neutral gas which is a non-supporter of combustion.
57. In the reaction - $2\text{KI} + \text{H}_2\text{O}_2 \rightarrow 2\text{KOH} + \text{I}_2$. Name the oxidizing agent & oxidized product.
58. A a] positive catalyst b] negative catalyst.
59. The type of chemical reaction seen in the chemical change - $\text{Cl}_2 + 2\text{KBr} \rightarrow 2\text{KCl} + \text{Br}_2$.
60. A trivalent metal which reacts with steam liberating hydrogen.
61. The shell or energy level having a maximum of 8 electrons.
62. The atom which needs one electron to attain stable electronic configuration of the nearest noble gas - helium.
63. A heterogeneous mixture of undissolved particles in the dispersion medium, existing in a state, too large to pass through a filter paper or a semi permeable membrane.