

ICSE Board
Class VII Chemistry
Sample Paper 1

Time: 2 hrs

Total Marks: 75

General Instructions:

1. *All questions are **compulsory**.*
 2. *Questions 1 to 15 carry one mark each.*
 3. *Questions in 2 A and B carry one mark each.*
 4. *Questions in 3 A and B carry one mark each.*
 5. *Questions in 4 A and B carry one mark each.*
 6. *Question 5 A and B carry five marks each.*
 7. *Questions in 6 A carry one mark each and 6B carry five marks.*
 8. *Question 7 carries ten marks in total.*
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Question 1

Choose the correct answer out of the four available choices given under each question. [15]

1. The _____ substances do not have their own definite shape and volume.
 - (a) crystal
 - (b) Solid
 - (c) Liquid
 - (d) Gaseous

2. Metals react with dilute acids to liberate _____ gas.
 - (a) Sulphur
 - (b) Oxygen
 - (c) Hydrogen
 - (d) Chlorine

3. A solution which cannot dissolve more of solute at a given temperature is called _____.
 - (a) Unsaturated solution
 - (b) Saturated solution
 - (c) True solution
 - (d) Colloid solution

4. Which of the following is a property of carbon dioxide
 - (a) Combustible, non-supporter of combustion
 - (b) Combustible, supporter of combustion
 - (c) Non-combustible, supporter of combustion
 - (d) Non-combustible, non-supporter of combustion

5. Phosphorus burns in oxygen to form
- (a) Phosphorus oxide
 - (b) Phosphorus dioxide
 - (c) Phosphorus tetroxide
 - (d) Phosphorus pentoxide
6. In a _____ reaction, two substances exchange their radicals and form a new substance.
- (a) Direct combination
 - (b) Decomposition
 - (c) Simple displacement
 - (d) Double displacement
7. Which method is based on the difference in weights of solid particles?
- (a) Filtration
 - (b) Sieving
 - (c) Handpicking
 - (d) Winnowing
8. Electrolysis of water is an example of _____ reaction.
- (a) Combination
 - (b) Decomposition
 - (c) Double decomposition
 - (d) Displacement
9. Burning of magnesium is a
- (a) Permanent
 - (b) Irreversible
 - (c) Chemical
 - (d) All of the above
10. In a chemical equation, the total mass of the reactants is equal to
- (a) Total quantity of the reactants
 - (b) Physical state of the reactants
 - (c) Total mass of the products
 - (d) Total mass of the solution
11. Sodium chloride is a
- (a) Acid
 - (b) Base
 - (c) Salt
 - (d) Alkali

12. The molecular formula of water is

- (a) OH
- (b) H₂O
- (c) H₂O₂
- (d) HO₂

13. The main chemicals that cause acid rain are _____

- (a) CO₂ and SO₂
- (b) SO₂ and NO₂
- (c) CO₂ and NO₂
- (d) O₂ and CO₂

14. The symbol for copper is

- (a) Co
- (b) Cp
- (c) Cu
- (d) Cr

15. Rusting of iron is a _____ change.

- (a) Chemical
- (b) Physical
- (c) Periodic
- (d) Fast

Question 2

(A) Give the chemical formula for the following: [5]

1. Caustic soda
2. Nitric acid
3. Lead nitrate
4. Blue vitriol
5. Vinegar

(B) Fill in the blanks and rewrite the sentences: [5]

1. Manganese dioxide acts as a _____ during the formation of oxygen.
2. On burning a candle, _____ and _____ are formed.
3. When ammonium chloride is dissolved in water, heat is _____.
4. A sublimable solid on heating turns into _____.
5. Adding salt to water increases the _____ point of water.

Question 3

(A) State whether True or False. [5]

1. Marble is a hydrated calcium sulphate.
2. Oxygen burns in an atmosphere of acetylene producing very high temperatures which is used for welding and cutting metals.
3. Metals occur in nature in both free and combined states.
4. Rusting is a slow oxidation process in which iron slowly reacts with oxygen of the air in the presence of moisture.
5. Water freezing into ice is an example of chemical change.

(B) Match the following: [5]

Solid 'X' to Liquid 'Y'	Liquefaction
Liquid 'X' to its vapours 'Z'	Solidification
'Z' to 'X'	Melting
'Y' to 'X'	Boiling point
The temperature at which 'Y' changes to 'Z'	Vaporisation

Question 4

(A) Write the valency and the symbols for the following elements: [5]

- a. Oxygen
- b. Sulphur
- c. Bromine
- d. Chlorine
- e. Carbon

(B) Give the chemical name for the following: [5]

1. NaOH
2. KOH
3. NaHCO₃
4. CaO
5. H₃PO₄

Question 5

(A) Give reasons: [5]

1. Addition of curd to milk is a chemical change.
2. Burning of sulphur powder is a chemical change.

(B) Explain the term catalyst and its types.

[5]

Question 6

(A) Define: [5]

1. Chemical equation
2. Reactants
3. Products
4. Balanced equation
5. Catalyst

(B) Compare in tabular form properties of Solids, Liquids and Gases with respect to : [5]

(i) Shape

(ii) Volume

(iii) Compressibility

(iv) Diffusion

(v) Fluidity or Rigidity

Question 7

1. Define mixture. What are its types? [4]
2. Explain the term atomicity in detail. [3]
3. Give the differences between elements and compounds. [3]

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Solution

Question 1

1. **(d)** Gaseous

The gaseous substances do not have their own definite shape and volume.

2. **(c)** Hydrogen

Metals react with dilute acids to liberate hydrogen gas.

3. **(b)** Saturated solution

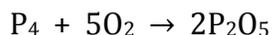
A solution which cannot dissolve more of a solute at a given temperature is called a saturated solution.

4. **(d)** Non-combustible, non-supporter of combustion

Carbon dioxide is non-combustible, non-supporter of combustion.

5. **(a)** Phosphorus pentoxide

Phosphorus burns in oxygen to form phosphorus pentoxide.



6. **(d)** Double displacement

In a double displacement reaction, two substances exchange their radicals and form a new substance.

7. **(d)** Winnowing

The winnowing method is based on the difference in weights of the solid particles.

Sieving and filtration methods are based on size and colour of the particles, and handpicking is based on the size or shape of the particles.

8. **(b)** Decomposition

Electrolysis of water is the decomposition of water (H₂O) into oxygen (O₂) and hydrogen (H₂) because of an electric current being passed through the water. This is an example of a decomposition reaction.

9. **(d)** All of the above

Burning of magnesium is a permanent, irreversible and chemical change.

10. (c) Total mass of the products

In a chemical equation, the total mass of the reactants is equal to

11. (c) Salt

Sodium chloride NaCl is a salt made of Na⁺ and Cl⁻ ions.

12. (b) H₂O

The molecular formula of water is H₂O.

13. (b) SO₂ and NO₂

The main chemicals that cause acid rain are SO₂ and NO₂.

14. (c) Cu

Copper is a chemical element with symbol Cu, from its Latin name **cuprum**.

15. (a) Chemical

When iron rusts, it reacts with moisture and oxygen present in the air to form a new compound called iron oxide (rust). This is a chemical change.

Question 2

(A)

1. Caustic soda: NaOH
2. Nitric acid: HNO₃
3. Lead nitrate: Pb(NO₃)₂
4. Blue vitriol: CuSO₄·5H₂O
5. Vinegar: CH₃COOH

(B)

1. Manganese dioxide acts as a catalyst during the formation of oxygen.
2. On burning a candle, carbon dioxide gas and water vapour are formed.
3. When ammonium chloride is dissolved in water, heat is absorbed.
4. A sublimable solid on heating turns into a gas.
5. Adding salt to water increases the boiling point of water.

Question 3

(A)

1. False. Gypsum is hydrated calcium sulphate.
2. True
3. True
4. True
5. False. False. Water freezing into ice is an example of physical change.

(B)

Solid 'X' to Liquid 'Y'	Melting
Liquid 'X' to its vapours 'Z'	Vaporisation
'Z' to 'X'	Solidification
'Z' to 'Y'	Condensation
The temperature at which 'Y' changes to 'Z'	Boiling point

Question 4

(A)

Element	Symbol	Valency
Oxygen	O	2
Sulphur	S	2
Bromine	Br	1
Chlorine	Cl	1
Carbon	C	4

(B)

1. Sodium hydroxide
2. Potassium hydroxide
3. Sodium bicarbonate
4. Calcium oxide
5. Phosphoric acid

Question 5

(A)

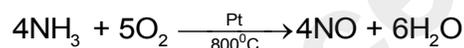
1. A chemical change is a change which is permanent and irreversible with no new substance formation. When curd is added to milk, the milk changes to curd and forms a new substance with altogether different composition and properties. This curd cannot be changed back into milk, and thus, this change is permanent and irreversible.
2. Burning of sulphur powder results in the formation of sulphur dioxide. This change is permanent and irreversible. Sulphur dioxide formed cannot be reversed into sulphur powder by altering the conditions. The composition and properties of sulphur dioxide is altogether different from those of sulphur powder. Thus, burning of sulphur powder is a chemical change.

(B)

A **catalyst** is a substance which alters (i.e. increases or decreases) the rate of the reaction without taking part in the reaction. It remains chemically unchanged at the end of the reaction.

Some reactions need a catalyst to speed up or slow down the reaction.

Example:



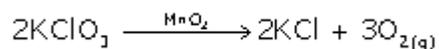
In this reaction ammonia and oxygen reacts in the presence of catalyst platinum at temperature 800°C to give nitric oxide and water.

Types of catalysts:

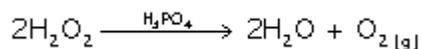
1. Positive catalyst: It is a catalyst which increases the rate of a chemical reaction.

Example:

Manganese oxide acts as a catalyst for the decomposition of potassium chlorate into potassium chloride and oxygen at a low temperature



2. Negative catalyst: It is a catalyst which decreases the rate of a reaction.



Question 6

(A)

1. Chemical equation: A chemical equation is a shorthand form representing the result of a chemical change.
2. Reactants: A substance that undergoes a chemical change during a reaction is called reactant..
3. Products: Products are substances which are formed as a result of a chemical reaction.
4. Balanced equation: A balanced equation is one in which the number of atoms of each element of the reactants is equal to the number of atoms of each element of the products.
5. Catalyst: A catalyst is a substance which speeds up a reaction but is chemically unchanged at the end of the reaction. After the reaction, the mass of the catalyst is exactly the same as that in the beginning.

(B)

Properties	Solids	Liquids	Gases
Shape	Definite shape	Do not have a definite shape, will take the shape of the container	No definite shape
Volume	Definite volume. As intermolecular forces between the constituent particles are strong	Definite volume. As intermolecular forces between the constituent particles are strong	No definite volume As intermolecular forces between the constituent particles are weak
Compressibility	Negligible	Negligible	High
Diffusion	Can diffuse into liquids	Diffusion is higher than solids	Highly diffusible as particles move randomly at high speed
Fluidity or rigidity	Very rigid and cannot flow from one place to another	Less rigid and are capable of flowing from higher to lower levels	No rigidity and can flow most easily among the three states of matter. They usually flow from high pressure to low pressure areas

Question 7

1. Mixtures contain two or more different substances mixed in any proportion.

- **Heterogeneous mixture**

A mixture in which the components or constituents are not uniformly distributed throughout its volume is called a heterogeneous mixture.

- The components or the constituents can be easily seen separately.
- Example: Oil in water

- **Homogeneous mixture**

A mixture in which the components or constituents are uniformly distributed throughout its volume is called a homogeneous mixture.

- The components or the constituents cannot be easily seen separately.
- Example: Salt in water

2. Atomicity

The number of atoms of an element which join to form a molecule of that element is known as atomicity of that molecule. Depending on the atomicity, the molecules of elements can be termed

- *Monatomic molecules*: Monatomic molecules contain only one atom.
Example: Zinc (Zn)
- *Diatomic molecules*: Diatomic molecules of an element contain two atoms of the same type. Example: Hydrogen molecule (H₂)
- *Triatomic molecules*: Triatomic molecules of an element contain three atoms of the same type. Example: Ozone (O₃)
- *Polyatomic molecules*: Polyatomic molecules of an element contain more than three atoms of the same type. Example: (P₄)

3.

Elements	Compounds
<ul style="list-style-type: none">• Elements are made of one kind of atoms.	<ul style="list-style-type: none">• Compounds are made of two or more kinds of atoms.
<ul style="list-style-type: none">• Elements cannot be broken down into simpler substances by any physical or chemical method.	<ul style="list-style-type: none">• Compounds can be broken down into simpler substances by chemical methods.
<ul style="list-style-type: none">• Elements have their own set of properties.	<ul style="list-style-type: none">• Properties of a compound differ from those of their elements.
<ul style="list-style-type: none">• Examples: Hydrogen, oxygen	<ul style="list-style-type: none">• Example: Water